In a neutralization reaction, a strong acid and a strong base react, and the acidic properties of the acid, as well as the basic properties of the base are canceled out. The products are a salt (which is an ionic compound made up of a metal and a nonmetal) and water (H₂O or HOH).

For each neutralization reaction, predict the products, circle the acid in the reactants, underline the salt in the products, and balance the equation.

1. $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H₂O}$

2. $\text{H₂SO₄} + \text{KOH} \rightarrow \text{K₂SO₄} + \text{H₂O}$

3. $\text{HNO₃} + \text{NaOH} \rightarrow \text{NaNO₃} + \text{H₂O}$

4. $\text{NaOH} + \text{HBr} \rightarrow \text{NaBr} + \text{H₂O}$

5. $\text{HI} + \text{KOH} \rightarrow \text{KI} + \text{H₂O}$

6. $\text{Ni(OH)₂} + \text{H₃PO₄} \rightarrow \text{Ni₃(PO₄)₂} + \text{H₂O}$